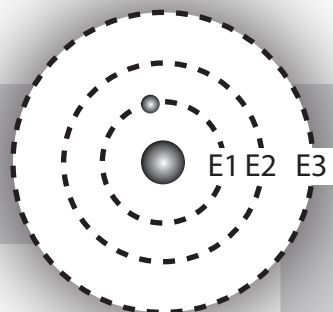
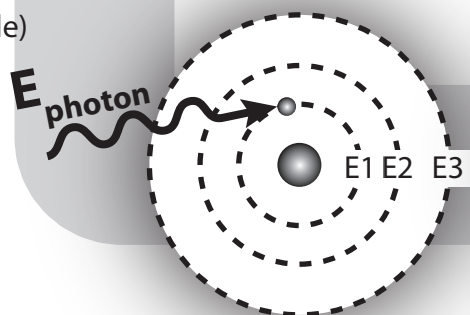


Quantum Leap



Ground State

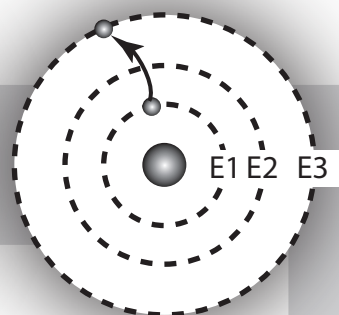
An electron's lowest possible energy level (E1 in this example)



Photon

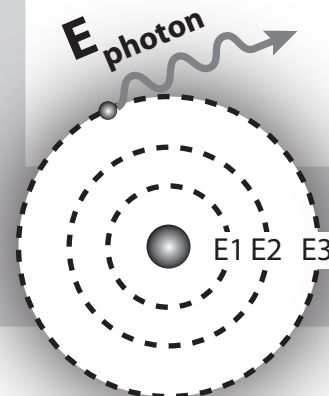
Absorption

The electron absorbs a photon with a particular amount of energy ($E_3 - E_1$)



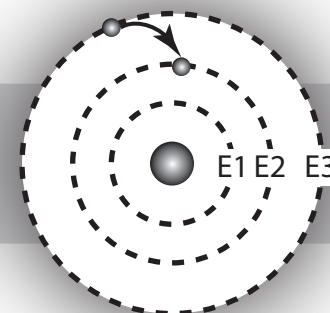
Excited State

The electron jumps to a higher energy level (E1 to E3)



Photon Emission

The electron emits a photon with a particular amount of energy ($E_3 - E_2$)



Return toward Ground State

The electron jumps to a lower energy level (E3 to E2)

Bohr Theory

Electrons can only orbit the nucleus of an atom in allowable paths or Energy Levels

Light

Both a photon and a wave

Quantum

A discrete quantity of something
A whole number of something...

Photon

A packet of energy